**Oxford High School**

**Chemistry**

**Chapter 1-7 Cumulative Review**

1. Black:white cation:\_\_\_\_\_\_\_\_\_
2. Anion
3. Sodium ion
4. Metal
5. Iron ion
6. Plant:oak tree anion:\_\_\_\_\_
7. Sodium ion
8. Cation
9. Magnesium ion
10. Chloride ion
11. Cl:Cl- Mg: \_\_\_\_\_
12. Mg2+
13. Mg2-
14. Al3+
15. Mn
16. Pipe:water metal:\_\_\_\_\_
17. Ions
18. Ionic bond
19. Electricity
20. Conductor
21. Consider the PI3 molecule
22. Draw its Lewis structure B. Draw the molecular shape including bond angles
23. How many protons, neutrons, and electrons are in
24. Zirconium -90 B. palladium-108 C. bromine-81
25. How is a molecule different from an atom?
26. Complete each of the following conversions (SHOW WORK)
27. 4.72 g to mg B. 2.7 x 103 cm/s to km/h C. 4.4 mm to dm
28. Name the following compounds
29. Fe(OH)3 B. NH4I C. Na2CO3 D. CCl4
30. Write the formula for the following compounds
31. Potassium nitrate B. copper II oxide C. magnesium nitride D. silver fluoride
32. Consider the selenium atom
33. Write its electron configuration
34. How many unpaired electrons are there?
35. What is the configuration for Se2-?
36. How many electrons are in the highest energy level for this ion?