Oxford High School

Chemistry

Chapter 1&2 Cumulative Review

1. The melting point of silver is 962oC. Express this temperature in Kelvins.
2. Classify each of the following as an element, compound, or mixture.
3. Egg
4. Cake
5. Dry ice, CO2
6. Iron powder
7. Describe the law of conservation of mass in terms of a campfire.
8. List three physical properties of each of the following objects:
9. A glass
10. A soda
11. Ice cubes
12. Bubbles (from a soda)
13. Indentify the larger quantity in each pair of measurements
14. Centigram – milligram
15. Deciliter – kiloliter
16. Millisecond – microsecond
17. Cubic decimeter – milliliter
18. Micrometer – nanometer
19. How many significant figures are in each of the following measurements?
20. 5.12 g
21. 3.456 x 106 kg
22. 0.000078 dm3
23. 0.04504 mm
24. 985.50 K
25. 65.02 s
26. Round each number in problem #6 to two significant figures

A.

B.

C.

D.

E.

F.

1. 121g of sulfuric acid is added to 4.00 x 102 mL of water. The resulting volume is 437 mL. What is the density of the solution?
2. The density of a piece of wood is 0.20 g/mL. If the wood has a mass of 98.0 g and measures 4.4 cm wide and 3.5 cm deep, what is the height of the block?
3. The density of air is 1.20 g/mL. What is the mass of air in kg in a room with the dimensions

25.0 m by 15.0m by 4.0m?

1. Complete the relationships
2. Journey:route problem: \_\_\_\_\_\_\_\_
3. Unknown
4. Plan
5. Known
6. Calculate
7. Meter:100cm gram:\_\_\_\_\_
8. Kg
9. 100 cm
10. 1000mg
11. 100 kg